Chemistry Lab “Finding the Pressure of a Carbon Dioxide Release”

Purpose:

Procedure: **YOU MUST WEAR SAFETY GLASSES AT ALL TIMES DURING THIS EXPERIMENT.**  Measure out exactly 40 ml of HCl into a graduated cylinder. Pour into an Erlenmeyer flask. Put between 5.0 and 8.0 grams of sodium hydrogen carbonate (baking soda) into a balloon. Put the balloon over the lip of the Erlenmeyer flask and pour the baking soda into the acid in the flask. Measure the circumference of the balloon and the temperature of the remaining liquid. Pour the contents of the flask down the sink and clean all equipment. Do only 1 trial

Materials:

Observations:

Data; Circumference of balloon\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Temperature\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Moles of baking soda\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Moles of carbon dioxide\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Volume of balloon\_\_\_\_\_\_\_\_\_\_\_\_\_

Conclusions: Pressure of the carbon dioxide\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_